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Model 2: Postulates of the Kinetic Molecular Theory POSTULATES OF THE KINETIC MOLECULAR THEORY 1. Gases consist of tiny particles (atoms or molecules). 2. These particles are so small compared with the distance between them that the volume (size) of the individual particles can be assumed to be negligible (zero). 3.

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TES OF THE KINETIC MOLECULAR THEORY Gases consist of tiny particles (atoms or molecules), These particles are so small, compared with the distances between them that the volume (size) of the individual particles can be assumed to be negligible (zero). The particles are in constant random motion, colliding with the walls of the container.

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Kinetic Molecular Theory Key Questions 1. What causes a gas to exert pressure when confined in a container? 2. How does the total volume of gas particles compare to the volume of the space between the gas particles? betwun 15 arc.) 3. As the temperature of a gas decreases, what change occurs in the amount of kinetic energy? Uaeases 4.

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The Kinetic-Molecular Theory Explains the Behavior of Gases, Part II According to Graham's law, the molecules of a gas are in rapid motion and the molecules themselves are small. The average distance between the molecules of a gas is large compared to the size of the molecules.

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